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representation of a chemical reaction
with reactants on the left, products
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Chapter 11 Chemical Reactions 113
SECTION 11.1 DESCRIBING CHEMICAL
REACTIONS (pages 321–329) This
section explains how to write
equations describing chemical

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Reactions using appropriate symbols.

It also describes how to write balanced chemical equations when given the names or formulas of the reactants and products in a chemical reaction. Writing Chemical Equations (pages 321–323) 1.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329)

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Chemical reaction with reactants on the left, products on the right, and an arrow separating the two skeleton
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Section 11.1 Assessment. Describe
the steps in writing a balanced

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chemical equation. Write the skeleton equation for the following reactions:
Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas.
Iron metal and chlorine gas react to form solid iron(III)chloride. $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$

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Reactions119 GUIDED PRACTICE

PROBLEM 2 (page 324) 2. Sulfur burns
in oxygen to form sulfur dioxide.

Write a skeleton equation for this
chemical reaction. Include

appropriate symbols from table 11.1.

Analyze Step 1. Identify the relevant
concepts. Write the formula for each
reactant and ...

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Ch. 11: Chemical Reactions

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Chemical Reactions. In a chemical reaction, the reactants are written on the left and the products on the right. The arrow that separates them is called yield.

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CHEMICAL REACTIONS. Chapter 11

Chemical Reactions 113. SECTION

11.1 DESCRIBING CHEMICAL

REACTIONS (pages 321–329). This

section explains GUIDED PRACTICE

PROBLEM 17 (page 334). 17.

Complete the equations for these single- replacement reactions in aqueous solution. Balance each

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because of chemical reactions.

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The initial step in solving a problem
of this type must be to write the
balanced chemical equation for the
reaction. Inspection of Equation
11.4.1 shows that it is balanced as

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written, so we can proceed to the
strategy outlined in Figure 11.4.1,
adapting it as follows:

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